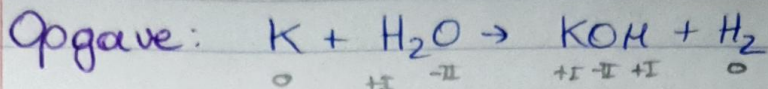
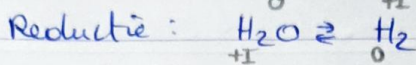
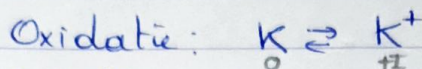


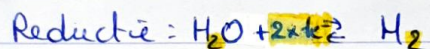
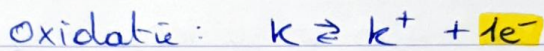
Methode: ion - elektronmethode



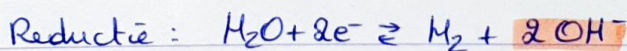
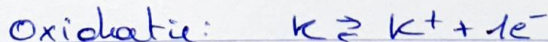
Stap 1: Zoek de redoxkoppels, duid de oxidatiegetallen aan en splits op in een oxidatie- en een reductie-deelreactie.



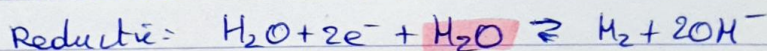
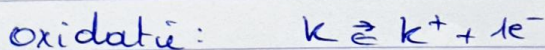
Stap 2: Zorg ervoor dat van de elementen die van OG veranderen, evenveel atomen voor als na de reactie voorkomen. Bepaal uit de verandering van de OG het aantal uitgewisselde elektronen.



Stap 3: Voeg in zuur milieu indien nodig H_3O^+ -ionen toe om in elke deelreactie te voldoen aan de wet van behoud van lading. Voeg in basisch milieu indien nodig OH^- -ionen toe.

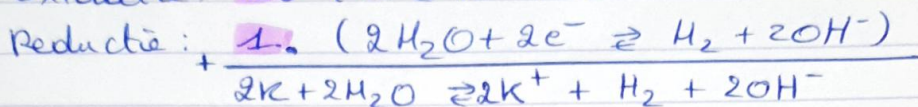


Stap 4: Voeg indien nodig water toe om in elke deelreactie te voldoen aan de wet van behoud van massa.

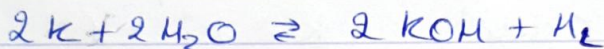


Stap 5: Het aantal afgestane elektronen moet gelijk zijn aan het aantal opgenomen elektronen. Vermenigvuldig indien nodig elke deelreactie met een zo klein mogelijk getal om hieraan te voldoen. Tel beide deelreacties op.

Na eventuele vereenvoudiging tot de kleinste mogelijke gehele getallen levert dit de essentiële deeltjesvergelijking.

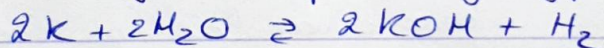


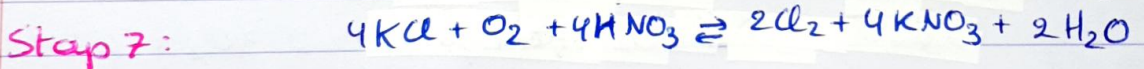
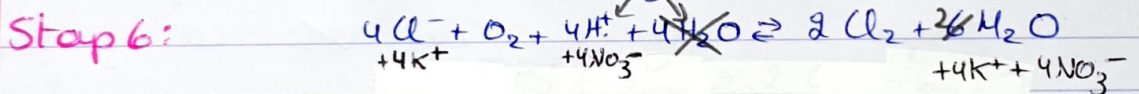
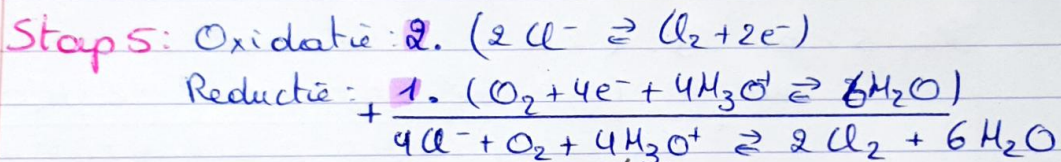
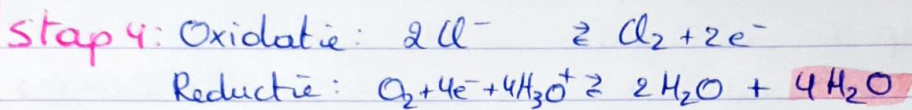
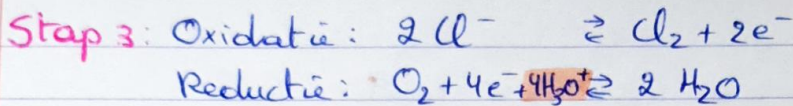
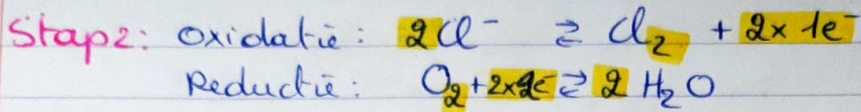
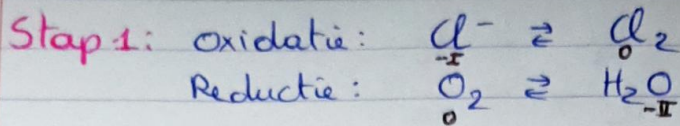
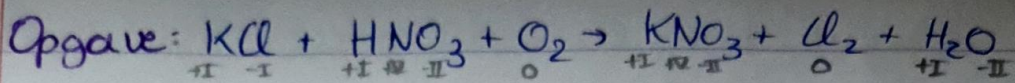
Stap 6: Vorm de stoffenreactie vergelijking door de toevoeging van de tegenionen. Controleer de atoombalans.

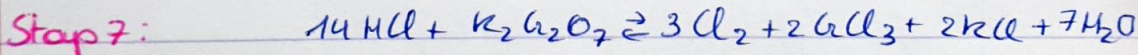
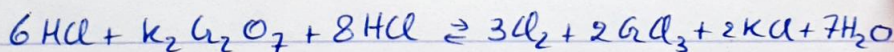
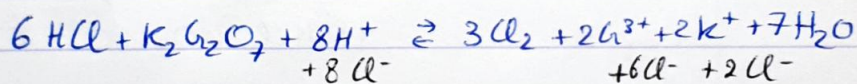
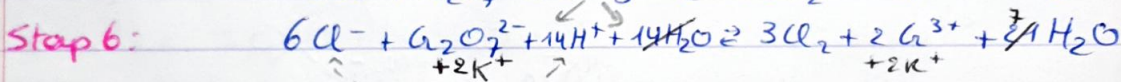
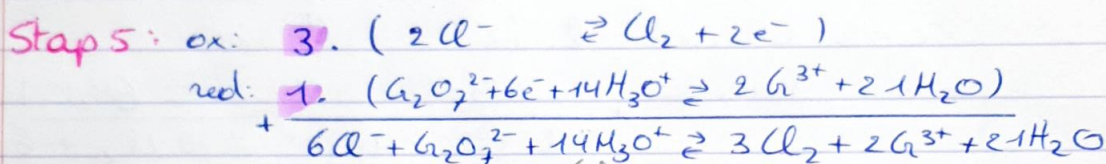
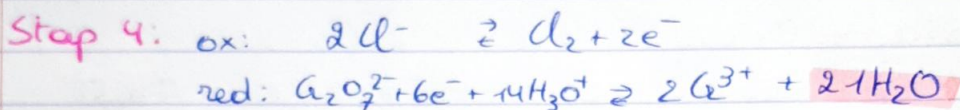
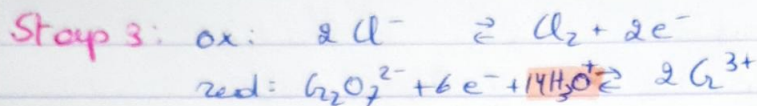
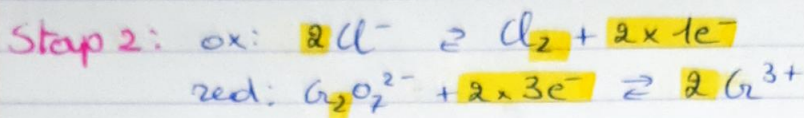
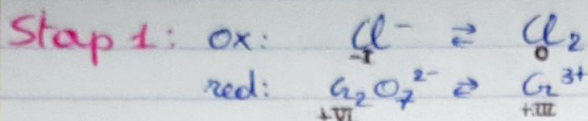
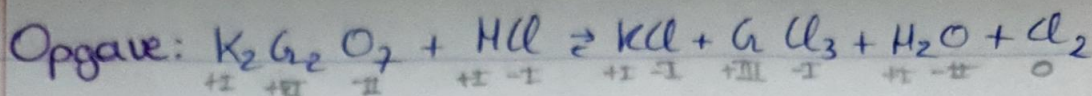


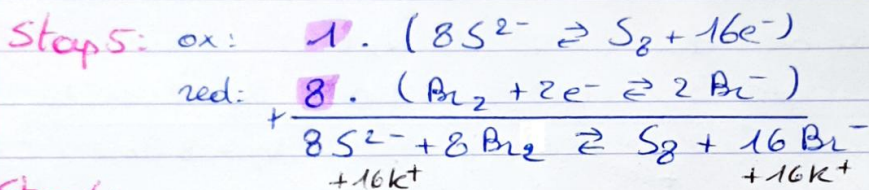
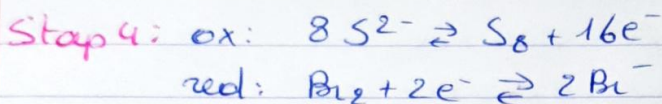
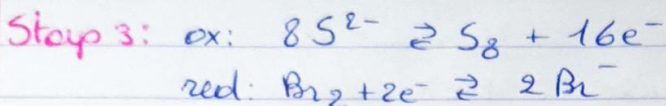
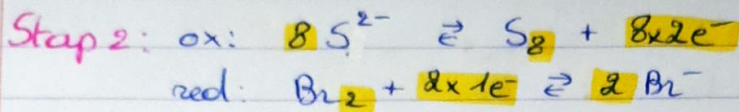
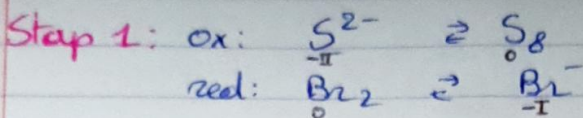
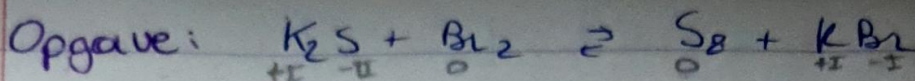
Stap 7: Vereenvoudig indien nodig.

Hier niet nodig. Onze oplossing blijft:

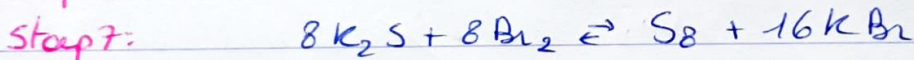


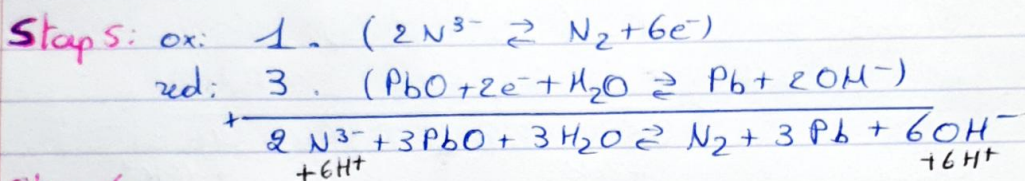
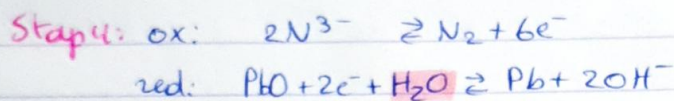
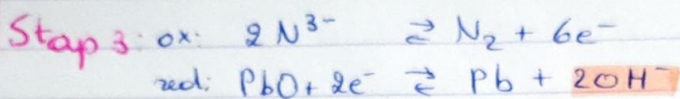
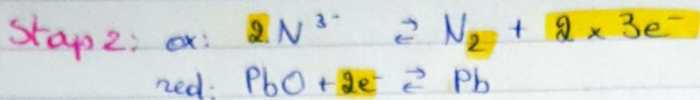
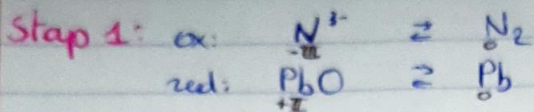
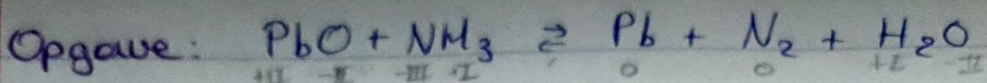






Step 6:





Step 6:

